

The Meaning of 4 Beryllium Being Followed By 9 Fluorine in the Periodic Table Rather Than 6 Carbon

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Abstract: In the periodic table of the elements 4 beryllium (no isotopes) holds the 2nd even position. The 3rd and 4th even positions, 6 carbon and 8 oxygen, are skipped in favor of the odd position 9 fluorine; why?

	stable isotopes	abundance(s)	duplications	density
1 hydrogen	1, 2	99.98%, 0.02	g/cm ³	0.07
2 helium	3, 4	0.0002%, 99.9998		0.14
3 lithium	6, 7	7.5%, 92.5		0.53
4 beryllium	9	100%,		1.85
5 boron	10, 11	20%, 80.0		2.08
6 carbon	12, 13	98.7%, 1.1		2.27
7 nitrogen	14, 15	99.6%, 0.4		0.80
8 oxygen	16, 17, 18	99.76%, 0.04, 0.20		1.14
9 fluorine	19	100%		1.70
10 neon	20, 21, 22	90.4%, 0.27, 9.25		1.21
11 sodium	23	100%		0.96
12 magnesium	24, 25, 26	79.0%, 10.0, 11.0,		1.73
13 aluminium	27	100%		2.70
14 silicon	28, 29, 30	92.2%, 4.7, 3.1		2.33
15 phosphorus	31	100%		2.34
16 sulfur	32, 33, 34, 36	94.99%, 0.75, 4.25, 0.01		1.84
17 chlorine	35, 37	76%, 24		1.56
18 argon	36 , 38, 40	0.33%, 0.06, 99.6		1.39
19 potassium	39, 41	93.25%, 6.73		0.86
20 calcium	40 , 42, 43, 44, 46	96.9%, 0.65, 0.135, 2.09, 0.004		1.55
21 scandium	45	100%		2.98
22 titanium	46 , 47, 48, 49, 50	8.2%, 7.4, 73.7, 5.3, 5.18		4.51

23 vanadium	51	99.75%	6.11
24 chromium	50 , 52, 53, 54	4.34%, 83.7, 9.50, 2.36	7.19
25 manganese	55	100%	7.21
26 iron	54 , 56, 57, 58	5.85%, 91.75, 2.12, 0.28	7.87
27 cobalt	59	100%	8.90
28 nickel	58 , 60, 61, 62, 64	68.1%, 26.2, 1.14, 3.63, 0.93	8.90

Starting with the nickel isotopes 60 and 62 which are even and the strongest bound of the periodic table, we have 4 no-isotope odd elements and then potassium and chlorine odd elements of 2 elements each followed again (starting at phosphorus) by 4 no-isotope odd elements ending with 9 fluorine. Taking (see my previous ViXra paper 2007.0229) odd periodic table elements for all stable elements up to 83, we get 4 groupings of 4 elements each: 19, 23, 27, 31 and 45, 51, 55, 59 and 75 (arsenic), **83** (rubidium), 89, 93 and 103, 113, 127, 133 then 1 grouping of 8 elements: 139 (lanthanum), 141, 153, 159, 165, 169, 175, 185 then finally 1 grouping of 4 elements: 197 (gold), **203**, 205 (thallium), 209 (bismuth - see Wikipedia), end of stable isotopes. We note that the groupings are alerting us first to the importance of MHCE8S theory (4 groups of 4 elements), then life (1 group of 8 elements) and finally the universe (1 group of 4 elements).

Returning to the question at hand, why are the positions 6 carbon and 8 oxygen skipped? - the answer is simple; **nature** does not want elements basic to life to be classified in the same group as elements such as fluorine (decorative etchant for glass) or cobalt (blue decorative color) or manganese and vanadium (alloys to make better swords for war), etc. On the other hand, in more recent years many of the group, phosphorus (fertilizer) and aluminum (airplanes) for example have become highly beneficial but only after science has been applied to improve the materials.