

Electron-Proton Asymmetry Hypothesis

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Summary

Electron-proton structure is already conclusive. It is pretty stupid to rethink about it. But there are always some things in my mind which I can't get rid of, so I can't help thinking about it for decades, you guys must understand that is really painful. I rather metaphorize proton and electron into male and female, more than into (+) and (-). Because in natural fact, there is no negative, and the proton meets the electron will not disappear, so they shall not mark as mathematics symbols. They attract each other as male and female. Now comes up the questions, how about they are not equal, like protons have more energy, or electrons do. How about they haven't got the notes from god, that they only can pair one another. After all, electrons equal protons and they match one on one, that is a hypothesis we put on them, but is that the fact? Let's do some hypotheses without this one.

Hypothesis

- Stability

Stability is more important than how we thought. In the universe, there are many different kinds of energy. Energies match others randomly. If the combination of energies is stable, they will keep staying in this firm, or they will separate then meet other energies. That model is more reasonable than human society. Stability is a very important guideline. Stabler combinations displace the less stable ones, that is the essence of chemical reactions. Apparently, electron-proton structure is stable.

- Electrons

Electrons don't take much space, we can imagine it has no volume. But between electrons, they are repulsive. Between the same kind of energy, they always are repulsive, or they will merge into a bigger unit. In the distance, they will kick each other away, then they won't touch each other. This distance makes an electron look like a sphere to other electrons, like pomegranate. But the size of electrons is only meaningful to electrons, and it is not fixed. If people use different particles to detect electrons, they will probably get different results, because different energies have different repulsive forces. In the atom, if the Coulomb force is larger, that will make the electrons stay closer, electrons close to nucleu will stay smaller than the outside ones.

- Protons

Protons are repulsive too. They are similar to electrons, but have different power and different energy types. Protons and electrons attract each other, the force is called Coulomb force. Protons and electrons can be spatial overlap.

- Nucleu

Nucleus is constituted by protons and electrons. If the Coulomb force is larger than the repulsive force, protons can be connected by electrons. The Coulomb force and repulsive force of protons will merge, same as electrons. Different forces don't offset, they don't work as the electron is negative then offset the proton's. Different forces stay in the same time and same space. Around nucleus, there is a strong Coulomb-Force-Field attracting electrons, and there is a different Repulsive-Force-Field to push electrons away from the nucleus. To a nucleus itself, how many protons doesn't really matter, but how stability does. We can find out in the element table, most elements have even mass, that is because even-protons structure is more stable than odd-protons, and odd-protons nucleus is transferred to even-protons in an instant after it exists.

- Repulsive-Force-Field

This is the most weird part of my hypothesis.

RFF blocks the electrons out from nucleus. When Coulomb force is larger than RFF, some electrons still can go through the RFF. The RFF will send these electrons far away, like the sorcerer's goal keepers always like to kick the soccer ball far away. It is not a real kicking, I rather call it teleport, directly teleporting the electrons from RFF to out of the atom. This action needs to consume some space energy, these space energy distinctions cause gravity.

- Atomic structure

The nucleus is the power resource of the atom. CFF attracts electrons to nucleus, and RFF blocks electrons from nucleus. Electrons and protons are asymmetry, even if there is only one proton in the nucleus, the CFF attract multiple electrons. The electrons fill the CFF until the CFF is not strong enough to hold more electrons.

As the number of protons goes up, the CFF and RFF go stronger and larger. The electrons look like covering the nucleus layer by layer. The electrons have no orbits (such as S-1, P-1). They simply move to the center of the atom, straightly. Because there are other electrons blocking in the front of the nucleus, the electrons move as they are squirming, they try every way to get into the nucleus.

With some number of protons, the RFF intensity is stronger than the CFF intensity, then there are barely electrons that get into CFF. Even if some outside electrons get in the CFF, the RFF still can block it. That looks like the atom would like to accept electrons and keep stable. They are known as Halogen.

With some number of protons, the RFF intensity and the CFF intensity reach kind of balance. The electrons fill the layer tight. The layer has no more space for any electron, and the CFF is not strong enough to catch one more layer. The RFF is strong enough to block any electron out. This perfect balance makes the atom look like they wouldn't get any outside electrons, and it wouldn't give out electrons either. They are known as Noble Gas.

With some number of protons, the RFF intensity is weaker than the CFF intensity, then there are some electrons drawn into RFF frequently. The electrons get into the RFF are teleported out of the atom quickly. Because the electrons are less, CFF have more space to catch outside electrons, it will fill the electron gap as soon as possible. So these kinds of atoms seem like they keep sending out electrons. They are known as Metal.

- Molecular structure

When Metal atoms meet the Halogen atoms, the Halogen keeps the electrons which Metal sends out. The Metal's CFF abstract surrounds electrons, and it holds the surface electrons of Halogen. These electrons are held by both atoms, and are known as Covalent bonds. Because of the Covalent bond, two atoms are bonded together.

Because the Halogen gets more electrons, electrons intensity is larger than RFF, a few electrons get into RFF and are sent out. These electrons are caught by the metal CFF as well.

The halogen and metal atom are bonded together, and stay in kind of moving balance.

This combination will appear attitudes like halogen, nobel gas and metal. They will abstract other atoms' electrons as covalent bonds. If that is more stable, they will keep this combination.